

SUPPORTING INFORMATION

Table TS1. Comparison of H₂ evolution from g-C₃N₄ based hybrid structures in the literature and present work.

g-C₃N₄ Morphology	Sacrificial agent	Wavelength (nm)	Lamp power (W)	Hydrogen evolution	Reference
TiO ₂ /Pt/g-C ₃ N ₄	TEOA	≥420	300	178 μmol.h ⁻¹	S1
Strontium pyroniobate/g-C ₃ N ₄	Methanol	420	300	57.5 μmol.h ⁻¹	S2
Ni(OH) ₂ -CdS/g-C ₃ N ₄	Na ₂ S/Na ₂ SO ₃	420	300	115.18 μmol.g ⁻¹ .h ⁻¹	S3
g-PAN/g-C ₃ N ₄	TEOA	>400	300	37 μmol.h ⁻¹	S4
Cd _{0.2} Zn _{0.8} S/g-C ₃ N ₄	Na ₂ S/Na ₂ SO ₃	>420	300	208 μmol.h ⁻¹	S5
MgFe ₂ O ₄ /g-C ₃ N ₄	TEOA	>420	300	30.09 μmol.h ⁻¹	S6
Ag ₂ O/g-C ₃ N ₄	TEOA	>420	300	~34 μmol.h ⁻¹	S7
CaIn ₂ S ₄ /g-C ₃ N ₄	Na ₂ S/Na ₂ SO ₃	420 nm	UV LEDs	102 μmol.g ⁻¹ .h ⁻¹	S8
CdS/Au/g-C ₃ N ₄	Methanol	>420	Not provided	19.02 μmol.g ⁻¹ .h ⁻¹	S9
TiO ₂ /g-C ₃ N ₄	NaI	365	8.3 mW/cm ²	45.6 μmol.h ⁻¹	S10
Ni(OH) ₂ /g-C ₃ N ₄	TEOA	≥420	300	240 μmol.g ⁻¹ .h ⁻¹	S11
ZnTiO ₃	TEOA	≥400	300	56.72 μmol.g ⁻¹	Present work
ZNTCN20				120.09 μmol.g ⁻¹	
ZNTCN60				295.88 μmol.g ⁻¹	
g-C ₃ N ₄				106.22 μmol.g ⁻¹	

Table TS1

Table TS2. Kinetic rate constant data of all pollutants, pure ZnTiO₃, pure g-C₃N₄, ZnTiO₃/g-C₃N₄ with 20 and 60 wt% of g-C₃N₄ visible light irradiation toward degradation of MB, phenol, 4-chlorophenol and 4-nitrophenol pollutants. Rate constant for corresponding solution without photocatalysts are also provided for comparison.

Sample details	Kinetic rate constant x 10 ⁻³ min ⁻¹			
	Methylene blue	Phenol	4-chlorophenol	4-nitrophenol
Without catalyst	0.23	0.31	0.46	0.32
ZnTiO ₃	2.51	0.95	0.91	2.83
g-C ₃ N ₄	2.69	2.29	3.06	9.91
ZnTiO ₃ /g-C ₃ N ₄ 20wt%	3.62	10.17	7.04	15.32
ZnTiO ₃ /g-C ₃ N ₄ 60wt%	8.53	34.37	17.39	30.64
P-25 (TiO ₂)	2.45	1.25	0.81	2.08

Table TS2

Figure S1 The kinetic plots for ZnTiO₃, g-C₃N₄ and ZnTiO₃/g-C₃N₄ samples are provided in (E) MB, (F) Phenol, (G) 4-chlorophenol and (H) 4-nitrophenol. The photocatalytic activity of commercial P-25 (TiO₂) powder is given for comparison.

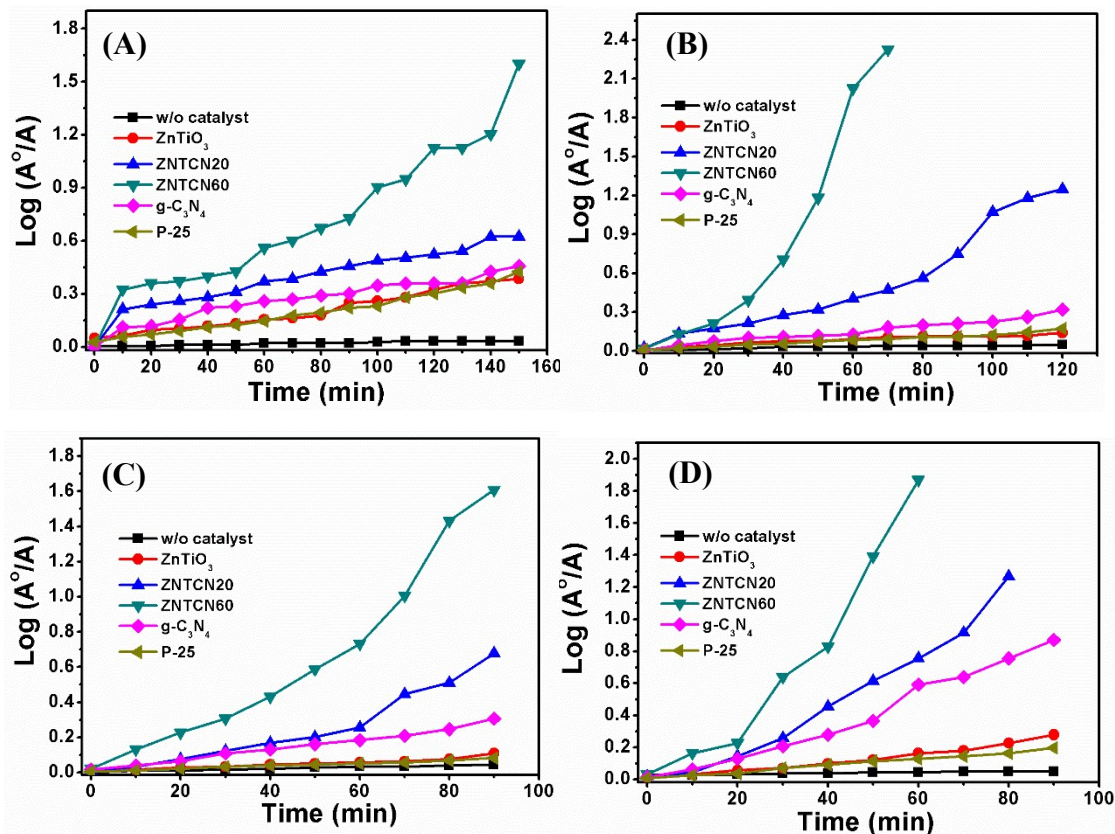


Figure S1 (A to D)

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